

BITSAT 2026 Chemistry Syllabus

1. States of Matter

1.1 Measurement: Physical quantities and SI units, Dimensional analysis, Precision, Significant figures.

1.2 Chemical reactions: Laws of chemical combination, Dalton's atomic theory; Mole concept; Atomic, molecular and molar masses; Percentage composition empirical & molecular formula; Balanced chemical equations & stoichiometry

1.3 Three states of matter, intermolecular interactions, types of bonding, melting and boiling points

Gaseous state: Kinetic energy and molecular speeds; Gas Laws, ideal behavior, ideal gas equation,

empirical derivation of gas equation, Avogadro number, Deviation from ideal behaviour – Critical

temperature, Liquefaction of gases, van der Waals equation.

1.4 Liquid state: Vapour pressure, surface tension, viscosity.

1.5 Solid state: Classification; Space lattices & crystal systems; Unit cell in two dimensional and three

dimensional lattices, calculation of density of unit cell – Cubic & hexagonal systems; Close packing;

Crystal structures: Simple AB and AB₂ type ionic crystals, covalent crystals – diamond & graphite,

metals. Voids, number of atoms per unit cell in a cubic unit cell, Imperfections - Point defects, non -

stoichiometric crystals; Electrical, magnetic and dielectric properties; Amorphous solids –

qualitative description. Band theory of metals, conductors, semiconductors and insulators, and n - and p - type semiconductors.

2. Atomic Structure

2.1 Introduction: Subatomic particles; Atomic number, isotopes and isobars, Thompson's model and

its limitations, Rutherford's picture of atom and its limitations; Hydrogen atom spectrum and Bohr model and its limitations.

2.2 Quantum mechanics: Wave -particle duality – de Broglie relation, Uncertainty principle; Hydrogen

atom: Quantum numbers and wave functions, atomic orbitals and their shapes (s, p, and d), Spin quantum number.

IV 2.3 Many electron atoms: Pauli exclusion principle; Aufbau principle and the electronic configuration of atoms, Hund's rule.

3. Periodicity and Bonding

3.1 Brief history of the development of periodic tables Periodic law and the modern periodic

table; Types of elements: s, p, d, and f blocks; Periodic trends: ionization energy, atomic, and ionic radii, inert gas radii, electron affinity, electronegativity and valency.

Nomenclature of elements with atomic number greater than 100 .

3.2 Valence electrons, Ionic Bond: Lattice Energy and Born -Haber cycle; Covalent character of ionic

bonds and polar character of covalent bond, bond parameters

3.3 Molecular Structure: Lewis picture & resonance structures, VSEPR model & molecular shapes

3.4 Covalent Bond: Valence Bond Theory - Orbital overlap, Directionality of bonds & hybridisation (s, p

& d orbitals only), Resonance; Molecular orbital theory - Methodology, Orbital energy level

diagram, Bond order, Magnetic properties for homonuclear diatomic species (qualitative idea only).

3.5 Dipole moments; Hydrogen Bond.

4. Thermodynamics

4.1 Basic Concepts: Systems and surroundings; State functions; Intensive & Extensive Properties;

Zeroth Law and Temperature

4.2 First Law of Thermodynamics: Work, internal energy, heat, enthalpy, heat capacities and specific

heats, measurements of ΔU and ΔH , Enthalpies of formation, phase transformation, ionization,

electron gain; Thermochemistry; Hess's Law, Enthalpy of bond dissociation, combustion, atomization, sublimation, solution and dilution

4.3 Second Law: Spontaneous and reversible processes; entropy; Gibbs free energy related to

spontaneity and non-spontaneity, non-mechanical work; Standard free energies of formation, free

energy change and chemical equilibrium

4.4 Third Law: Introduction

5. Physical and Chemical Equilibria

5.1 Concentration Units: Mole Fraction, Molarity, and Molality

5.2 Solutions: Solubility of solids and gases in liquids, Vapour Pressure, Raoult's law, Relative lowering

of vapor pressure, depression in freezing point; elevation in boiling point; osmotic pressure,

determination of molecular mass; solid solutions, abnormal molecular mass, van't Hoff factor.

Equilibrium: Dynamic nature of equilibrium, law of mass action.

5.3 Physical Equilibrium: Equilibria involving physical changes (solid-liquid, liquid-gas, solid-gas).

5.4 Chemical Equilibria: Equilibrium constants (K_P , K_C), Factors affecting equilibrium, Le-Chatelier's

principle.

5.5 Ionic Equilibria: Strong and Weak electrolytes, Acids and Bases (Arrhenius, Lewis, Lowry and

Bronsted) and their dissociation; degree of ionization, Ionization of Water; ionization of polybasic

acids, pH; Buffer solutions; Henderson equation, Acid-base titrations; Hydrolysis; Solubility Product

of Sparingly Soluble Salts; Common Ion Effect.

5.6 Factors Affecting Equilibria: Concentration, Temperature, Pressure, Catalysts, Significance of ΔG and

ΔG° in Chemical Equilibria.

6. Electrochemistry

6.1 Redox Reactions: Oxidation -reduction reactions (electron transfer concept);
Oxidation number;
Balancing of redox reactions; Electrochemical cells and cell reactions; Standard electrode potentials;
EMF of Galvanic cells; Nernst equation; Factors affecting the electrode potential;
Gibbs energy change and cell potential; Secondary cells; dry cells, Fuel cells; Corrosion and its prevention.
V 6.2 Electrolytic Conduction: Electrolytic Conductance; Specific and molar conductivities; variations of conductivity with concentration, Kolhrausch's Law and its application, Electrolysis, Faraday's laws of electrolysis; Electrode potential and electrolysis.

7. Chemical Kinetics

7.1 Aspects of Kinetics: Rate and Rate expression of a reaction; Rate constant; Order and molecularity of the reaction; Integrated rate expressions and half-life for zero and first order reactions.
7.2 Factor Affecting the Rate of the Reactions: Concentration of the reactants, catalyst; size of particles, Temperature dependence of rate constant concept of collision theory (elementary idea, no mathematical treatment); Activation energy, Arrhenius Equation .

7.3 Surface Chemistry: Adsorption – physisorption and chemisorption; factors affecting adsorption of gasses on solids; catalysis: homogeneous and heterogeneous, activity and selectivity: enzyme catalysis, colloidal state: distinction between true solutions, colloids and suspensions; lyophilic, lyophobic multi molecular and macromolecular colloids; properties of colloids; Tyndall effect, Brownian movement, electrophoresis, coagulations; emulsions –types of emulsions.

8. Hydrogen and s-block elements

8.1 Hydrogen: Element: unique position in periodic table, occurrence, isotopes;

Dihydrogen:

preparation, properties, reactions, and uses; Molecular, saline, ionic, covalent, interstitial hydrides;

Water: Properties; Structure and aggregation of water molecules; Heavy water;

Hydrogen peroxide:

preparation, reaction, structure & use, Hydrogen as a fuel.

8.2 s-block elements: Abundance and occurrence; Anomalous properties of the first elements in each group; diagonal relationships; trends in the variation of properties (ionization energy, atomic &

ionic radii).

8.3 Alkali metals: Lithium, sodium and potassium: occurrence, extraction, reactivity, and electrode

potentials; Biological importance; Reactions with oxygen, hydrogen, halogens and water; Basic

nature of oxides and hydroxides; Halides; Properties and uses of compounds such as

NaCl, Na₂CO₃,

NaHCO₃, NaOH, KCl, and KOH.

8.4 Alkaline earth metals: Magnesium and calcium: Occurrence, extraction, reactivity and electrode

potentials; Reactions with O₂, H₂O, H₂ and halogens; Solubility and thermal stability of oxo salts;

Biological importance of Ca and Mg; Preparation, properties and uses of important compounds

such as CaO, Ca(OH)₂, plaster of Paris, MgSO₄, MgCl₂, CaCO₃, and CaSO₄.

9. p- d- and f-block elements

9.1 General: Abundance, distribution, physical and chemical properties, isolation and uses of elements;

Trends in chemical reactivity of elements of a group; electronic configuration, oxidation states;

anomalous properties of first element of each group.

9.2 Group 13 elements: Boron; Properties and uses of borax, boric acid, boron hydrides & halides.

Reaction of aluminum with acids and alkalis;

9.3 Group 14 elements: Carbon: carbon catenation, physical & chemical properties, uses, allotropes

(graphite, diamond, fullerenes), oxides, halides and sulphides, carbides; Silicon: Silica, silicates,

silicone, silicon tetrachloride, Zeolites, and their uses

9.4 Group 15 elements: Dinitrogen; Preparation, reactivity and uses of nitrogen; Industrial and

biological nitrogen fixation; Compound of nitrogen; Ammonia: Haber's process, properties and

reactions; Oxides of nitrogen and their structures; Properties and Ostwald's process of nitric acid

production; Fertilizers – NPK type; Production of phosphorus; Allotropes of phosphorus; Preparation, structure and properties of hydrides, oxides, oxoacids (elementary idea only) and

halides of phosphorus, phosphine.

9.5 Group 16 elements: Isolation and chemical reactivity of dioxygen; Acidic, basic and amphoteric

oxides; Preparation, structure and properties of ozone; Allotropes of sulphur;

Preparation/production properties and uses of sulphur dioxide and sulphuric acid; Structure and

properties of oxides, oxoacids (structures only).

VI 9.6 Group 17 and group 18 elements: Structure and properties of hydrides, oxides, oxoacids of

halogens (structures only); preparation, properties & uses of chlorine & HCl; Inter halogen

compounds; Bleaching Powder; Uses of Group 18 elements, Preparation, structure and reactions of

xenon fluorides, oxides, and oxoacids.

9.7 d-Block elements: General trends in the chemistry of first row transition elements; Metallic

character; Oxidation state; ionization enthalpy; Ionic radii; Color; Catalytic properties; Magnetic

properties; Interstitial compounds; Occurrence and extraction of iron, copper, silver, zinc, and

mercury; Alloy formation; Steel and some important alloys; preparation and properties of $K_2Cr_2O_7$,

$KMnO_4$.

9.8 f-Block elements: Lanthanoids and actinoids ; Oxidation states and chemical reactivity of

lanthanoids compounds; Lanthanide contraction and its consequences, Comparison of actinoids

and lanthanoids.

9.9 Coordination Compounds: Coordination number; Ligands; Werner's coordination theory; IUPAC

nomenclature; Application and importance of coordination compounds (in qualitative analysis,

extraction of metals and biological systems e.g. chlorophyll, vitamin B12 , and hemoglobin);

Bonding: Valence -bond approach, Crystal field theory (qualitative); Isomerism including stereoisomerisms.

10. Principles of Organic Chemistry and Hydrocarbons

10.1 Classification: General Introduction, classification based on functional groups, trivial and IUPAC

nomenclature. Methods of purification: qualitative and quantitative,

10.2 Electronic displacement in a covalent bond: Inductive, resonance effects, and hyperconjugation;

free radicals; carbocations, carbanions, nucleophiles and electrophiles; types of organic reactions,

free radical halogenations.

10.3 Alkanes: Structural isomerism, general properties and chemical reactions, free radical halogenation, combustion and pyrolysis.

10.4 Alkenes and alkynes: Structure of double and triple bonded compounds; General methods of

preparation and reactions, physical properties, electrophilic and free radical additions, addition of

hydrogen, halogen, water, hydrogen halides; (Markovnikov's addition and peroxide effect),

ozonolysis, oxidation, mechanism of electrophilic addition; acidic character of alkynes and (1,2 and

1,4) addition to dienes.

10.5 Aromatic hydrocarbons: Sources; properties; isomerism; resonance delocalization; aromaticity;

polynuclear hydrocarbons; IUPAC nomenclature; mechanism of electrophilic substitution reaction,

directive influence and effect of substituents on reactivity; carcinogenicity and toxicity.

10.6 Haloalkanes and haloarenes: Physical properties, nomenclature, optical rotation, chemical reactions

and mechanism of substitution reaction. Uses and environmental effects; di, tri, tetrachloromethanes, iodoform, freon and DDT.

11. Stereochemistry

11.1 Conformations: Ethane conformations; Newman and Sawhorse projections.

11.2 Geometrical isomerism in alkenes

12. Organic Compounds with Functional Groups Containing Oxygen and Nitrogen

12.1 General: Nomenclature, electronic structure, important methods of preparation, identification,

important reactions, physical and chemical properties, uses of alcohols, phenols, ethers, aldehydes,

ketones, carboxylic acids, nitro compounds, amines, diazonium salts, cyanides and isocyanides.

12.2 Specific: Reactivity of alpha -hydrogen in carbonyl compounds, effect of substituents on alpha -

carbon on acid strength, comparative reactivity of acid derivatives, mechanism of nucleophilic

addition and dehydration, basic character of amines, methods of preparation, and their separation,

importance of diazonium salts in synthetic organic chemistry.

VII 13. Biological, Industrial and Environmental chemistry

13.1 Carbohydrates: Classification; Monosaccharides; Structures of pentoses and hexoses; Simple

chemical reactions of glucose, Disaccharides: reducing and non -reducing sugars – sucrose, maltose

and lactose; Polysaccharides: elementary idea of structures of starch, cellulose and glycogen.

13.2 Proteins: Amino acids; Peptide bond; Polypeptides; Primary structure of proteins; Simple idea of secondary, tertiary and quaternary structures of proteins; Denaturation of proteins and enzymes.

13.3 Nucleic Acids: Types of nucleic acids; Primary building blocks of nucleic acids (chemical

composition of DNA & RNA); Primary structure of DNA and its double helix.

13.4 Vitamins: Classification, structure, functions in biosystems; Hormones.

13.5 Polymers: Classification of polymers; General methods of polymerization; Molecular mass of polymers; Biopolymers and biodegradable polymers; methods of polymerization (free radical,

cationic and anionic addition polymerizations); Copolymerization: Natural rubber; Vulcanization of

rubber; Synthetic rubbers. Condensation polymers. Some important polymers: natural and

synthetic like polythene, nylon, polyesters, bakelite, and rubber.

13.6 Pollution: Environmental pollutants; soil, water and air pollution; Chemical reactions in atmosphere;

Smog; Major atmospheric pollutants; Acid rain; Ozone and its reactions; Depletion of ozone layer

and its effects; Industrial air pollution; Greenhouse effect and global warming; Green Chemistry,

study for control of environmental pollution.

13.7 Chemicals in medicine, health-care and food: Analgesics, Tranquilizers, antiseptics, disinfectants,

anti-microbials, anti-fertility drugs, antihistamines, antibiotics, antacids; Preservatives, artificial

sweetening agents, antioxidants, soaps and detergents.

14. Theoretical Principles of Experimental Chemistry

14.1 Volumetric Analysis: Principles; Standard solutions of sodium carbonate and oxalic acid; Acid - base

titrations; Redox reactions involving KI, H₂SO₄, Na₂SO₃, Na₂S₂O₃ and H₂S; Potassium permanganate

in acidic, basic and neutral media; Titrations of oxalic acid, ferrous ammonium sulphate with

KMnO₄, K₂Cr₂O₇/Na₂S₂O₃, Cu(II)/Na₂S₂O₃.

14.2 Qualitative analysis of Inorganic Salts: Principles in the determination of the cations Pb²⁺, Cu²⁺,

As³⁺, Mn²⁺, Al³⁺, Zn²⁺, Co²⁺, Ca²⁺, Sr²⁺, Ba²⁺, Mg²⁺, NH⁴⁺, Fe³⁺, Ni²⁺ and the anions CO₃²⁻, S²⁻, SO₄²⁻

, SO₃²⁻, NO₂⁻, NO₃⁻, Cl⁻, Br⁻, I⁻, PO₄³⁻, CH₃COO⁻, C₂O₄²⁻.

14.3 Physical Chemistry Experiments: preparation and crystallization of alum, copper sulphate. Benzoic

acid ferrous sulphate, double salt of alum and ferrous sulphate, potassium ferric sulphate;

Temperature vs. solubility; Study of pH changes by common ion effect in case of weak acids and

weak bases; pH measurements of some solutions obtained from fruit juices, solutions of known

and varied concentrations of acids, bases and salts using pH paper or universal indicator; Lyophilic

and lyophobic sols; Dialysis; Role of emulsifying agents in emulsification. Equilibrium studies

involving ferric and thiocyanate ions (ii) [Co(H₂O)₆]²⁺ and chloride ions; Enthalpy determination for

strong acid vs. strong base neutralization reaction (ii) hydrogen bonding interaction between

acetone and chloroform; Rates of the reaction between (i) sodium thiosulphate and hydrochloric

acid, (ii) potassium iodate and sodium sulphite (iii) iodide vs. hydrogen peroxide, concentration

and temperature effects in these reactions.

14.4 Purification Methods: Filtration, crystallization, sublimation, distillation, differential extraction, and

chromatography. Principles of melting point and boiling point determination; principles of paper

chromatographic separation – Rf values.

14.5 Qualitative Analysis of Organic Compounds: Detection of nitrogen, sulphur, phosphorous and

halogens; Detection of carbohydrates, fats and proteins in foodstuff; Detection of alcoholic,

phenolic, aldehydic, ketonic, carboxylic, amino groups and unsaturation .

VIII 14.6 Principles of Organic Chemistry Experiments: Preparation of, acetanilide, p-nitro acetanilide, di-

benzyl acetone, aniline yellow, beta -naphthol -aniline dye.

14.7 Basic Laboratory Technique: Cutting glass tube and glass rod, bending a glass tube, drawing out

a glass jet, boring of cork.